Unit 2: Generating and Burning Hydrogen Gas Lab

OBSERVATIONS:

1. Describe the chemical changes that occurred to the zinc metal.

* It started bubbling.

1. How do you know that there was hydrogen gas in the large test tube? Explain.

* It started fogging up.

QUESTIONS / CONCLUSION:

1. Calculate the mass difference in materials after the experiment*. Ask the teacher to help you with this.* If there isn’t much difference, explain why.

A: Empty test tube plus rack: 319.99

B: Full test tube plus rack: 381.89

C: Test tube plus rack:

B-A = 61.9 Mass of reactants

C-A = 19.6 Mass of products

1. How can the difference in mass be explained? How might we “conserve” the mass? Explain.

Because there was a chemical reaction.

1. We are going to explore this reaction type a bit later in the unit. For now, please write down the chemical formula (symbol and charge) for Zinc, HCl and Hydrogen Gas.

Zinc: ZN 2

HCL: Hydrochloric acid 0

Hydrogen gas: H 0